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(54) Title: CATALYTIC COMPOSITION SUITABLE FOR THE FISCHER-TROPSCH PROCESS (57) Abstract Catalytic composition which can be used in the Fischer-Tropsch process essentially consisting of an inert carrier, cobalt in a quantity ranging from 1 to 50 % by weight, preferably from 5 to 35 % by weight, and tantalum in a quantity ranging from 0.05 % to 5 % by weight, preferably from 0.1 to 3 % by weight.		

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CATALYTIC COMPOSITION SUITABLE FOR THE FISCHER-TROPSCH
PROCESS

The present invention relates to a catalytic composition suitable for the preparation reaction of
5 hydrocarbons by means of the so-called Fischer-Tropsch synthesis; it also relates to the catalytic process for the preparation of hydrocarbons for which it is used.

More specifically, the present invention relates to a new catalytic composition for the production of
10 hydrocarbons by means of the Fischer-Tropsch synthesis containing Cobalt and promoted by Tantalum, this composition being obtained by reacting derivatives of the above elements in the presence of a suitable carrier, as specified in more detail hereunder.

15 The selection of cobalt as main constituent of the active phase is due to the fact that this favours the formation of saturated linear hydrocarbons with a high molecular weight minimizing the formation of oxygenated and olefinic compounds, contrary to the well-known

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catalytic systems based on iron.

The known art cites numerous examples of catalysts based on cobalt used for the synthesis of paraffinic products with various distributions.

5 Since the first works of Fischer in 1932 (H.H. Storch, N. Golumbic, R.B. Anderson, "The Fischer Tropsch and Related Synthesis", John Wiley & son, Inc., New York, 1951) - which described the development of a Co/ThO₂/MgO system supported on kieselguhr - until
10 today, the patented systems based on cobalt are essentially the following: Co/Mg/ThO₂ supported on kieselguhr (1954, Reinpruessen A.G.), Co/MgO supported on bentonite (1958, M.W. Kellog), Co/Th/Mg (1959, Rurchemie), Co/Th supported on silica-gel (1960, Esso Res. & Eng.),
15 Co/Mg/Zr/Kieselguhr (1968, SU-A-660.324, Zelinskii INST.), Co/Ru/Keiselguhr (1976, US-A-4.088.671 GULF), Co/Zr/SiO₂ (1980, GB-A-2.073.237, Shell), Co/Ru supported on titanium (1988, US-A-4.738.948 Exxon), Co/Re/REO,K supported on alumina (1988, EP-A-313.375,
20 Statoil), Co/Mo,W/K,Na/SiO₂ (1991, GB-A-2.258.414, IFP), Co/Ru/Cu/K,Sr/SiO₂ (1993, EP-A-581.619, IFP).

The effect of promoters on the system based on cobalt, from what is described in literature, is multiple; however it can be subdivided into various groups
25 in relation to the function of the promoter (B. Jager,

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R. Espinoza in Catalysis Today 23, 1995, 21-22).

For example, promoters such as K, Na, Mg, Sr, Cu, Mo, W and metals of group VIII essentially increase the activity. Ru, Zr, rare-earth oxides (REO), Ti increase
5 the selectivity to hydrocarbons with a high molecular weight. Ru, REO, Re, Hf, Ce, U, Th favour the regenerability of the catalyst.

Among the various promoters, ruthenium, alone or together with other elements, is certainly the most
10 widely used.

The recent evolution of catalytic systems for the synthesis of hydrocarbons has led to the identification of various promoters to be coupled with cobalt in order to increase both the activity of these systems in terms
15 of conversion of the reagents and also the selectivity to linear hydrocarbons with a high molecular weight. This evolution has taken place mainly in the last twenty years. The increase in price of crude oil in the 70s' provided the incentive for exploring other ways of
20 producing liquid fuels and chemicals, among which the possibility of transforming synthesis gases into hydrocarbon products with a high molecular weight by means of the Fischer-Tropsch synthesis.

As far as the Fischer-Tropsch synthesis is concerned, this can refer to the hydrogenation process of
25

carbon monoxide to produce higher hydrocarbons and oxygenated molecules with a prevalently linear chain. The reaction takes place in the presence of a mixture of hydrogen and carbon monoxide with or without carbon
5 dioxide (so-called synthesis gas) at temperatures lower than 350°C and at pressures of between 1 and 100 atm.

The wide range of catalysts and their modifications described in the known art and the wide range of operating conditions for the reduction reaction of
10 carbon monoxide with hydrogen allows a considerable flexibility in the selectivity of the products, ranging from methane to heavy waxes with alcohols and olefins as by-products. The distribution of the products can be explained by the known growth mechanism obtained by a
15 polymerization kinetics and processed by Anderson Shultz and Flory (P. Biloen, W.M.H. Sachtler, Advance in Catalysis, Vol. 30, pages 169-171, Academic Press, New York, 1981; R.B. Anderson, Catalysis, Vol. IV, P.H. Emmett ed., Reinhold, New York, 1956). In accordance
20 with this model, the attempt to restrict the range of products in order to maximize for example the C₅-C₁₁ fraction (gasoline-range) results in selectivities to methane and the C₂-C₄ fraction of more than 40%. In addition the products obtained are essentially paraf-
25 fins with a linear chain and olefins with a low octane

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number. The only possibility of deviating from the nature imposed by the Fischer-Tropsch polymerization kinetics is to identify catalytic systems which do not adhere to this kinetics mechanism. Typical examples are
5 systems developed by Mobil which substantially couple the properties of the Fischer-Tropsch catalysts with the shape selectivity of zeolites (US-A-4.157.338).

The possibility of maximizing the selectivity to heavy liquids and waxes (essentially paraffinic and
10 without sulfur) offers on the other hand various advantages. In particular it is possible to minimize the selectivity to methane and the gas fraction. The subsequent treatment (e.g. hydrocracking, hydroisomerization) of this liquid-solid fraction of a paraf-
15 finic nature gives high quality medium distillates if compared with the medium distillates obtained from petroleum (Ball J., Gas. Matters, April 27 1989, pages 1-8). In this context the typical characteristic of catalysts based on cobalt to be highly selective in the
20 production of higher paraffins is definitely advantageous. In addition, the use of catalysts with a reduced water gas shift activity, such as catalysts based on cobalt, implies a low selectivity to CO₂, contrary to the use of a traditional catalyst based on iron.

25 With respect to the productivity of catalysts

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based on cobalt, defined as weight of C_{2+} hydrocarbons/weight of catalyst/time, from what is specified in literature, this seems to vary a great deal and to depend directly on the operating temperature. Increasing the operating temperature however is not a valid way of increasing the productivity to high quality liquid and solid hydrocarbons, as this would cause a consequent increase in the selectivity to methane and light gases. It is very important on the contrary from an economical point of view to maximize this productivity and to minimize at the same time the selectivity to methane. In other words it is important to maximize the production of high quality liquid and solid hydrocarbons (C_{9+} , C_{22+}).

15 In accordance with this important objective, it is necessary for the catalyst to be capable of combining a high productivity (Prod. C_{2+}) with a low selectivity to methane (Sel. CH_4).

A catalytic composition has now been found which, applied to the Fischer-Tropsch process, enables a high selectivity to C_{2+} hydrocarbons and at the same time a low selectivity to methane.

In accordance with this, the present invention relates to a catalytic composition based on cobalt which allows conversions of the mixture of CO and H_2 ,

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known as synthesis gas, with or without N_2 and/or CO_2 and/or light gases (C_1 - C_4), into saturated linear hydrocarbons containing from 77% to 88% by weight of C_{5+} and from 24 to 31% by weight of C_{9+} and with productivities into C_{2+} of between 180 and 330 $gC_{2+}/Kg_{cat}/h$, maintaining a low selectivity to methane.

The catalytic composition of the present invention essentially consists of an inert carrier, cobalt in a quantity of from 1 to 50% by weight, preferably from 5 to 35% by weight, and tantalum in a quantity of from 0.05 to 5% by weight, preferably from 0.1 to 3% by weight, the complement to 100 consisting of the inert carrier, the cobalt and tantalum being present in metal form or in the form of a derivative.

The percentages of cobalt and tantalum are expressed as metals.

The cobalt and tantalum can be present as metal or as derivatives, in the latter case the oxide form being preferred.

As far as the inert carrier is concerned, this is preferably selected from at least one of the oxides of at least one of the following elements: silicon, aluminum, zinc, magnesium, titanium, zirconium, yttrium, tin and the relative mixtures.

The inert carrier which can be used is independent

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of the crystallographic structure of the above oxides. For example, aluminas can be used of any phase composition such as η , γ , δ , θ , α and the relative mixtures.

In the same way, when the inert carrier essentially consists of TiO_2 , this can be in the form of rutile and/or anatase.

In the preferred embodiment, the inert carrier is selected from silica, γ alumina, δ alumina, titania and the relative mixtures, even more preferably from silica, γ alumina and the relative mixtures.

A further object of the present invention relates to a process for the preparation of the catalytic composition of the present invention which comprises:

- a) a first deposition onto the inert carrier, preferably selected from silica and alumina, preferably via dry impregnation, of a cobalt salt; subsequent calcination to give a catalytic precursor; subsequent optional reduction and passivation of the calcined product;
- b) deposition onto the catalytic precursor thus obtained of a derivative of tantalum, preferably via wet impregnation; subsequent calcination, optionally followed by reduction and passivation.

The cobalt and tantalum can be deposited according to various methods well known to experts in the field such as, for example, ion exchange, dry impregnation,

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wet impregnation; precipitation and coprecipitation, gelation and mechanical mixing.

In the case of cobalt however, the dry impregnation method is preferable. According to this technique, 5 the material to be impregnated is put in contact with a volume of solution more or less equal to the pore volume. In step (a) it is preferable to use aqueous solutions of cobalt salts, such as halides, nitrate, oxalate, the complex formed with lactic acid and 10 lactates, the complex formed with tartaric acid and tartrates, the complex formed with acetylacetonates. In the most preferred embodiment, cobalt nitrate is used.

In the case of tantalum on the other hand, this is preferably deposited by means of any impregnation 15 technique, preferably wet impregnation. According to this technique, the inert carrier onto which the cobalt has been previously deposited, is completely covered with a solution of a derivative of tantalum, particularly tantalum alcoholates, such as ethoxide, propoxi- 20 de, isopropoxide, methoxide. In the most preferred embodiment tantalum ethoxide dissolved in C₁-C₅ alcohols is used.

The inert carrier can be used partially or totally in the first phase. In the latter case all of the inert 25 carrier is used in the preparation of the catalytic

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precursor in the first step. In the former case the inert carrier is used partially in the first step and partially in the second step.

In the preferred embodiment, the process of the present invention comprises the above steps a) and b) without the reduction and passivation phases.

As far as the calcination is concerned, this is a heating step at a temperature of between 400°C and 750°C to remove volatile substances and decompose the derivatives of cobalt and tantalum into oxides. The calcination is carried out in the presence of oxygen, air or other gases containing oxygen.

Before this step, the material can be subjected to drying, usually at reduced pressure at a temperature of between 80 and 120°C, with or without an inert gas. This operation has the purpose of eliminating (or strongly reducing) the possible solvents or water with which the material has been impregnated and gives dispersion homogeneity to the active phase.

With respect to the reduction, this is a step in which the material is treated with a reducing agent, preferably hydrogen or gas containing hydrogen. The reduction is carried out at a temperature of between about 250°C and about 500°C, preferably from 300° to 450°C for periods of time of between 0.5 and 24 hours,

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at pressures between atmospheric pressure and 40 bars.

As far as the passivation is concerned, this is carried out by treating the material with oxygen diluted with an inert gas, usually nitrogen. The temperature is preferably from 10 to 80°C. Using for example nitrogen containing 1-2% of oxygen with a stream of 2 litres/hour/g_{cat.}, the passivation step can have a duration of from 1 to 5 hours at 25°C.

Some of the operating details relating to the preparation of the above catalytic compositions will however be more evident on reading the experimental examples below which, however, do not limit the catalytic compositions of the present invention.

A further object of the present invention relates to a process for the preparation of essentially linear, saturated hydrocarbons starting from synthesis gas (Fischer-Tropsch process) in the presence of the above catalytic composition.

The conversion of the synthesis gas into hydrocarbons takes place at a pressure normally between 1 and 100 bars, preferably from 10 to 75 bars, at a temperature generally within the range of 150°C to 350°C, preferably from 170°C to 300°C, even more preferably from 200°C to 240°C. The hourly volumetric flow-rate is generally from 100 to 20000, preferably from 400 to

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5000, volumes of synthesis gas per volume of catalyst and per hour. The ratio H_2/CO in the synthesis gas is generally from 1:2 to 5:1, preferably from 1.2:1 to 2.5:1. Other gases, particularly CO_2 may also be present.

As is known to experts in the field, the preparation of the mixture of CO and H_2 can be carried out starting from natural gas prevalently consisting of methane. The oxidating agent can be oxygen or air. In the latter case it is evident that the mixture of synthesis gas will also contain a considerable quantity of nitrogen, which may or may not be eliminated of CO/H_2 before the Fischer-Tropsch reaction. The advantage of carrying out the Fischer-Tropsch reaction on mixtures in which nitrogen is still present, is evident.

The catalyst can be used in the form of fine powder (about 10-700 mm) or in the form of particles having an equivalent diameter of from 0.7 to 10 mm, respectively in the presence of a liquid phase (under the operating conditions) and a gaseous phase, or a gaseous phase. The liquid phase can consist of at least one hydrocarbon having at least 5, preferably at least 15, carbon atoms per molecule. In the preferred embodiment, the liquid phase essentially consists of the same reaction product.

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Just to give an example, the catalysts of the present invention can be used in a fixed-bed reactor, fed in continuous with a mixture of CO and H₂ and operating under the following conditions:

- 5 - reaction temperature: 200-240°C
- reaction pressure: 20 bars
- space velocity (GHSV): 500-1500 h⁻¹
- H₂/CO mixture: 2/1

The reaction temperature is regulated in order to
10 obtain a conversion higher than at least 45% of the volume of carbon monoxide fed (conv. CO%).

Following these conditions, the catalysts prepared as described in examples 1 to 11 were evaluated, using various carriers. The compositions are summarized in
15 table 1.

The results of the reactivity tests are indicated in tables 2 to 4.

Catalysts supported on SiO₂

Comparative example 1.

20 Comparative catalyst A

(Co/Ru/SiO₂; 14% Co, 0.2% Ru).

A silica carrier (having a surface area of 520 m²/g, a specific pore volume of 0.8 m³/g, an average particle diameter of 0.5 mm, a specific weight of 0.42
25 g/ml) is dry impregnated with a nitric solution of

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Co(NO₃)₂·6H₂O at pH = 2.5 in such quantities as to obtain a percentage of Co equal to 14% by weight referring to the total. The silica thus impregnated is dried at 120°C for 16 hours, calcined at 400°C in air for 4 hours, then treated in a stream of H₂ at a space velocity (GHSV) of 1000 h⁻¹, in a tubular reactor at 400°C for 16 hours. The sample thus reduced is passivated in a mixture of (1%)O₂/(99%)N₂ with GHSV of 1000 h⁻¹ for 2 hours at room temperature.

10 A 7.5 10⁻³ M solution is added to the monometallic sample, of Ru(NO₃)₃·xH₂O obtained with the following procedure: precipitation in the form of hydroxide at pH = 7.2 of RuCl₃·xH₂O, subsequent elimination of the chlorides, resolubilization in conc. HNO₃ and dilution
15 in CH₃COCH₃ in a ratio 1:250 v/v.

The acetone solution of ruthenium is added to the sample in such a quantity as to have 0.2% of Ru by weight referring to the total. The slurry is left under stirring for 2 hours and then dried under vacuum
20 < 10mmHg at 50°C. A calcination phase in air follows at 350°C for 4 hours and subsequently a reduction and passivation analogous to that described above.

Comparative example 2.

Comparative catalyst B

25 (Co/Sc/SiO₂; 14% Co, 0.2% Sc).

For the preparation of catalyst B, a solution of $\text{Sc}(\text{NO}_3)_3$ 10^{-3} M in acetone is added to 50 g of the monometallic catalyst Co/SiO_2 prepared as described in example 1, in such a volume as to obtain a final weight percentage of Sc equal to 0.2%.

The suspension thus obtained is left under stirring for two hours and is then dried under vacuum at 50°C . The sample is calcined at 350°C for 4 hours in air, reduced at 400°C in H_2 for 16 hours with a GHSV of 1000 h^{-1} and passivated in (1%) O_2 /(99%) N_2 with a GHSV of 1000 h^{-1} for 2 hours at room temperature.

Example 3. Catalyst C1

($\text{Co}/\text{Ta}/\text{SiO}_2$; 14% Co, 0.5% Ta).

A solution of $\text{Ta}(\text{EtO})_5$ 0.01 M in ethanol is added to 50 g of the monometallic catalyst Co/SiO_2 prepared as described in example 1, in such a volume as to obtain a final weight percentage of tantalum equal to 0.5%.

The suspension thus obtained is left under stirring for two hours and is then dried under vacuum at 50°C .

The sample is calcined at 350°C for 4 hours in air, reduced at 400°C in H_2 for 16 hours with a GHSV of 1000 h^{-1} and passivated in (1%) O_2 /(99%) N_2 with a GHSV of 1000 h^{-1} for 2 hours at room temperature.

Example 3b. Catalyst C2

(Co/Ta/SiO₂; 14% Co, 0.2% Ta).

The catalyst C2 is prepared analogously to what is described in example 3.

5 Example 4. Catalyst D

(Co/Ta/SiO₂; 14% Co, 0.5% Ta).

A silica carrier (having a surface area of 520 m²/g, a specific pore volume of 0.8 m³/g, an average particle diameter of 0.5 mm, a specific weight of 0.42 g/ml) is dry impregnated with a nitric solution of Co(NO₃)₂·6H₂O at pH = 2.5 in such quantities as to obtain a percentage of Co equal to 14% by weight referring to the total. The silica thus impregnated is dried at 120°C for 16 hours and calcined at 400°C in
10
15 air for 4 hours. A solution of Ta(EtO)₅ 0.01 M in ethanol is added to the monometallic sample Co/SiO₂ in such a volume as to obtain a final weight percentage of tantalum equal to 0.5%.

The suspension thus obtained is left under stirring for two hours and is then dried under vacuum at 50°C.
20

A calcination phase in air follows at 350°C for 4 hours.

Catalyst supported on TiO₂

25 Comparative example 5. Comparative catalyst E

(Co/Ru/TiO₂; 12% Co, 0.2% Ru).

Following the procedure described in example 1, comparative catalyst E is prepared completely similar to catalyst A but having TiO₂ as carrier instead of SiO₂. In this case the TiO₂ had a surface area of 25 m²/g, a specific pore volume of 0.31 cm³/g and a content of rutile equal to 81%.

Comparative example 6. Comparative catalyst F

(Co/Sc/TiO₂; 12% Co, 0.2% Sc)

10 Catalyst F is prepared analogously to what is described for the preparation of catalyst B.

Example 7. Catalyst G

(Co/Ta/TiO₂; 12% Co, 0.5% Ta).

Following the procedure described in example 4, catalyst G is prepared, consisting of the carrier based on titania. In this case the TiO₂ had a surface area of 25 m²/g, a specific pore volume of 0.31 cm³/g and a content of rutile equal to 81%.

Catalysts supported on SiO₂-TiO₂

20 **Example 8. Catalyst H**

(Co/Ta/[Si-Ti]; 15% Co, 0.5% Ta).

A silica carrier (having a surface area of 480 m²/g, a specific pore volume of 0.8 m³/g, a particle size of between 75 and 150 μm, a specific weight of 25 0.55 g/ml, average pore radius of 35 Å) previously

dried at 150°C for 8 hours is suspended, under a nitrogen atmosphere, in anhydriified n-hexane, 6 ml/g SiO₂. A solution of Ti(i-PrO)₄ 0.2 M is added to the slurry in such a quantity as to have about 7.0% of Ti; 5 the mixture is left under stirring for 16 hours and is then dried under vacuum with a pressure of <10 mmHg and a temperature of 50°C. The sample thus obtained is calcined in a nitrogen atmosphere at 400°C for 4 hours and subsequently calcined in air at 600°C for a further 10 4 hours.

Catalyst H is prepared with the mixed carrier thus obtained, consisting of 7.1% of titanium of which about 25% in crystalline form (50% rutile, 50% anatase) and a surface area of 440 m²/g, analogously to what is 15 described in example 4.

Catalysts supported on Al₂O₃

Example 9. Catalyst I

(Co/Ta/Al₂O₃; 14% Co, 0.5% Ta).

Catalyst I is prepared analogously to what is 20 described in example 4, with an alumina carrier (crystalline phase 100% gamma, surface area of 175 m²/g, specific pore volume 0.5 m³/g, average pore radius 40 Å, particle size between 20-150 μm, specific weight of 0.86 g/ml).

25 Example 10. Catalyst L

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(Co/Ta/Al₂O₃; 12% Co, 0.5% Ta).

Catalyst L is prepared analogously to what is described in example 4, with an alumina carrier (crystalline phase 50% γ and 50% δ , surface area of 137 m²/g, specific pore volume 0.46 m³/g, average pore radius 45 Å, particle size between 20-120 μ m, specific weight of 0.69 g/ml).

CATALYTIC TESTS

Example 11. Evaluation of the catalytic activity of catalysts supported on silica.

The catalyst (A, B, C, D according to examples 1-4) is formed in particles having a diameter of between 0.35 and 0.85 mm and is subsequently diluted with an inert carrier, silicon carbide, having the same particle size as the catalyst and in a volumetric ratio catalyst/inert carrier equal to 1:2. The catalyst thus diluted is then charged into a tubular reactor and subjected to an activation procedure in a stream of hydrogen (2000 Nl/h·l_{cat}) and nitrogen (1000 Nl/h·l_{cat}), at a temperature of between 350-400°C and a pressure of 1 bar for 16 hours. The temperature is then lowered to 180°C, the volumetric flow-rate of hydrogen and nitrogen is modified (333-1000 Nl/h·l_{cat}) and (5000-15000 Nl/h·l_{cat}) respectively, the system is pressurized to 20 bars and carbon monoxide is then introduced (116.5-500

Nl/h·l_{cat}) to obtain a volumetric ratio H₂/CO equal to 2.

The flow-rate of nitrogen in the starting phase of the reaction is progressively lowered until complete elimination according to the following sequence (the
5 lower flow-rates refer to tests with GHSV = 500h⁻¹, the higher flow-rates to GHSV = 1500 h⁻¹):

time	flow-rate H ₂	flow-rate CO	flow-rate N ₂
(hrs)	(Nl/h·l _{cat})	(Nl/h·l _{cat})	(Nl/h·l _{cat})
0	333-1000	166.5-500	5000-15000
10 1	333-1000	166.5-500	3750-11250
2	333-1000	166.5-500	2500-7500
3	333-1000	166.5-500	1250-3750
4	333-1000	166.5-500	0

At the end of the starting phase, the reaction
15 temperature is regulated so as to obtain a conversion of carbon monoxide with respect to the volume fed (conv. CO%) of less than 20% for at least 48 hours, then in the following 48 hours the temperature is gradually increased until a minimum CO conversion value
20 of 45% is reached, without however exceeding the reaction temperature of 240°C, in order to minimize the production of methane as well as the light gaseous fractions (C₂-C₄).

As indicated in table 2 for comparative catalyst
25 A, in order to reach conversions of CO higher than the

limit of 45%, it is necessary to increase the reaction temperature (from 200°C to 240°C) with an increase in the volumetric flow-rates of the mixture H_2 -CO (GHSV from 500h⁻¹ to 1500h⁻¹). As a result the selectivity to methane is favoured (from 7.8% to 29.7%), expressed as percentage referring to the total carbon present in the products (C%), to the total disadvantage of the selectivities to higher hydrocarbons (Sel.C₂₂₊ from 15.4% to 3.2%, Sel.C₉₊ from 66.9% to 48.8%), expressed as percentage referring to the total weight of the whole hydrocarbon fraction produced (weight %).

With respect to comparative catalyst B, promoted with scandium, using a total volumetric flow-rate equal to 1500h⁻¹ and a reaction temperature of 218°C, an hourly weight productivity to hydrocarbons with more than two carbon atoms (C₂₊) equal to 273 g/kg/h and selectivities to C₂₂₊ of 14.2% are obtained. In general the catalytic performances of catalyst B can be considered as being higher than those of catalyst A.

Catalysts C1, C2 and D of the present invention, containing Tantalum, are subjected to an analogous catalytic test. As indicated in table 2, with a total volume flow-rate (GHSV) equal to 1500h⁻¹ and a reaction temperature of 220°C, for catalysts C1 and C2, prepared with the same procedure as catalysts A and B, CO

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conversions are obtained of 60.3% and 69.3% respectively, productivities to C_{2+} of more than 315 $gC_{2+}/Kg_{cat}/h$, selectivities to methane of less than 10%, selectivities to C_{22+} higher hydrocarbons of about 24%,
5 selectivities to C_{9+} hydrocarbons of between 65.6% and 71.3% and finally selectivities to C_{5+} of more than 81%.

These performances are better than those obtained with comparative catalysts A and B, especially for the higher productivities, selectivities to higher hydro-
10 carbons and lower selectivities to methane and light gas fractions (C_2-C_4).

With respect to catalyst D, synthesized with the procedure described in example 4, the catalytic performances of the Co/Ta system are further improved compared to the comparative catalysts: CO conversions of
15 71.0%, productivities to C_{2+} of 330 $gC_{2+}/Kg_{cat}/h$, selectivities to methane of 8.4%, selectivities to C_{22+} higher hydrocarbons of 29.1%, selectivities to C_{9+} hydrocarbons of 78.4% and finally selectivities to C_{5+}
20 of 83.5%.

Example 12. Evaluation of the catalytic activity of catalysts supported on Titania

As indicated in table 3, also in this case the comparison between the reference catalysts promoted
25 with ruthenium (cat. E) or scandium (cat. F) and the

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catalyst promoted without the intermediate reduction and passivation phase (example 3), shows an increase in the CO conversion, total hydrocarbon productivity and selectivity to higher hydrocarbons, maintaining a low
5 selectivity to methane (CO conv. = 70.0%, Prod.C₂₊ = 172 g/Kg_{cat}/h, C₂₂₊ = 32.9%, CH₄ = 7.6%).

Example 13. Evaluation of the catalytic activity of catalysts supported on silica/titania and alumina.

The catalytic composition Co/Ta supported on other
10 materials such as the mixed carrier silica-titania and alumina with a different phasic composition showed interesting catalytic performances at reaction temperatures of between 209 and 218°C and total volumetric flow-rates of 1500 h⁻¹.

15 As indicated in table 4, the conversions obtained are higher than 57% (CO conv. = 65.8÷57.1%), productivities to C₂₊ higher than 180 g/kg/h (Prod.C₂₊ : 183.1÷260.1 g/kg/h), selectivities to hydrocarbons C₂₂₊ higher than 23% (Sel.C₂₂₊: 23.2÷28.3).

20 The data of table 5 show the possibility of using synthesis gas diluted with nitrogen.

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Table 1

	Example	Cat.	%Co	%X	X Carrier
	Comp. 1	A	14	0.2	Ru SiO_2
5	Comp. 2	B	14	0.2	Sc SiO_2
	3	C1	14	0.5	Ta SiO_2
	3b	C2	14	0.2	Ta SiO_2
	4	D	14	0.5	Ta SiO_2
	Comp. 5	E	12	0.2	Ru TiO_2
10	Comp. 6	F	12	0.2	Sc TiO_2
	7	G	12	0.5	Ta TiO_2
	8	H	12	0.5	Ta $\text{SiO}_2\text{-TiO}_2$
	9	I	14	0.5	Ta $\gamma\text{-Al}_2\text{O}_3$
	10	L	12	0.5	Ta $\gamma, \delta\text{-Al}_2\text{O}_3$
15					

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Table 2

Example 11					
Catalysts supported on Silica					
5	Catalyst	A	A	A	B
	React. temp. (°C)	200	220	240	218
	GHSV h ⁻¹	500	1.000	1.500	1.500
	CO conv. (%)	48.5	51.3	47.1	57.6
	Prod.C ₂₊ (g/Kg/h)	81.4	149	183.5	272.6
10					
	CH ₄ (C%)	7.8	18.8	29.7	10.3
	CO ₂ (C%)	0.3	1.8	2.2	0.1
	C ₁ -C ₄ (wt%)	13.5	32.7	49	17.3
	C ₂₂₊ (wt%)	15.4	14.1	3.2	14.2
15	C ₉₊ (wt%)	66.9	64.1	48.8	-
	C ₅₊ (wt%)	86.5	67.3	51	82.8

Tabl 2 cont.

Example 11				
Catalysts supported on Silica				
5	Catalyst	C1	C2	D
	React. temp. (°C)	212	222	220
	GHSV h ⁻¹	1.500	1.500	1.500
	CO conv. (%)	60.3	69.3	71
	Prod. C ₂₊ (g/Kg/h)	315.3	326.2	330.1
10				
	CH ₄ (C%)	7.1	9.5	8.4
	CO ₂ (C%)	0.9	0.5	0.6
	C ₁ -C ₄ (wt%)	15.9	19	16.5
	C ₂₂₊ (wt%)	23.9	24.7	29.1
15	C ₉₊ (wt%)	65.6	71.3	78.4
	C ₅₊ (wt%)	84.1	81.1	83.5

Table 3

Example 12			
Catalysts supported on Titania			
5 Catalyst	E	F	G
React. temp. (°C)	228	228	214
GHSV h ⁻¹	1.500	1.500	1.500
CO conv. (%)	54.3	61.6	70
Prod. C ₂₊ (g/Kg/h)	144.6	152.3	171.7
10			
CH ₄ (C%)	7.6	13.3	7.6
CO ₂ (C%)	0.1	0.1	0.1
C ₁ -C ₄ (wt%)	12.6	19	12.1
C ₂₂₊ (wt%)	21.9	17.6	32.9
15 C ₅₊ (wt%)	87.4	81	87.9

Table 4. Catalysts supported on Si-Ti and alumina

Example 13				
Catalysts supported on Si-Ti and alumina				
5	Catalyst	H	I	L
	React. temp. (°C)	209	214	218
	GHSV h ⁻¹	1.500	1.500	1.500
	CO conv. (%)	65.8	59	57.1
	Prod. C ₂₊ (g/Kg/h)	260.1	183.1	196.2
10				
	CH ₄ (C%)	10.5	9.8	8.5
	CO ₂ (C%)	0.4	0.5	0.5
	C ₁ -C ₄ (wt%)	19.4	22.4	17.6
	C ₂₂₊ (wt%)	23.2	28.3	24.8
15	C ₉₊ (wt%)	68.2	65.5	70.4
	C ₅₊ (wt%)	80.6	77.6	82.4

Table 5

Test in the presence of Nitrogen			
5	Catalyst	I	I
	React. temp. (°C)	226	235
	GHSV h ⁻¹	1.500	1.500
	Content & N ₂	20.6	20.6
	CO conv. (%)	44	65.4
10	Prod. C ₂₊ (g/Kg/h)	120.4	173.5
	CH ₄ (C%)	15	13.4
	CO ₂ (C%)	0.4	0.9
	C ₁ -C ₄ (wt%)	27	28
15	C ₂₂₊ (wt%)	28.2	21.4
	C ₉₊ (wt%)	59.1	62.6
	C ₅₊ (wt%)	73	72

CLAIMS

1. A catalytic composition essentially consisting of an inert carrier, cobalt in a quantity of from 1 to 50% by weight, and tantalum in a quantity of
5 from 0.05 to 5% by weight,
the complement to 100 consisting of the inert carrier;
the cobalt and tantalum being present in metal form or in the form of a derivative.
- 10 2. The catalytic composition according to claim 1, characterized in that the cobalt is present in a quantity of from 5 to 35% by weight and the tantalum from 0.1 to 3% by weight.
3. The catalytic composition according to claim 1,
15 characterized in that the inert carrier is selected from at least one of the oxides of at least one of the following elements: silicon, aluminum, zinc, magnesium, titanium, zirconium, yttrium, tin and the relative mixtures.
- 20 4. The catalytic composition according to claim 1, characterized in that the inert carrier is selected from silica, γ alumina, δ alumina, titania and the relative mixtures.
5. The catalytic composition according to claim 4,
25 characterized in that the inert carrier is select-

ed from silica, γ -alumina and the relative mixtures.

6. A process for the preparation of the catalytic composition according to claim 1 which comprises:
 - 5 a) a first deposition onto the inert carrier, preferably selected from silica and alumina, of a cobalt salt; subsequent calcination to give a catalytic precursor; subsequent optional reduction and passivation of the calcined product;
 - 10 b) deposition onto the catalytic precursor thus obtained of a derivative of tantalum; subsequent calcination, optionally followed by reduction and passivation.
7. The process according to claim 6, characterized in
15 that the cobalt is deposited onto the inert carrier following the dry impregnation technique.
8. The process according to claim 6, characterized in that the tantalum is deposited via wet impregnation.
- 20 9. A process for the synthesis of essentially linear and saturated hydrocarbons starting from synthesis gas basically consisting of CO and H₂, possibly diluted with nitrogen, comprising reacting this mixture in the presence of the catalyst according
25 to claim 1, at a pressure of between 1 and 100

INTERNATIONAL SEARCH REPORT

Internal Application No

PCT/EP 98/04035

A. CLASSIFICATION OF SUBJECT MATTER

IPC 6 B01J23/847 C07C1/04 B01J23/755

According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

IPC 6 B01J C07C

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practical, search terms used)

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
A	US 4 328 158 A (INNES ROBERT A ET AL) 4 May 1982	
A	US 5 559 065 A (LAUTH GUENTER ET AL) 24 September 1996	
A	US 4 206 135 A (KUGLER EDWIN L ET AL) 3 June 1980	
A	US 4 738 948 A (FIATO ROCCO A ET AL) 19 April 1988 cited in the application	
A	US 4 088 671 A (KOBYLINSKI THADDEUS P) 9 May 1978 cited in the application	

☐ Further documents are listed in the continuation of box C.

☒ Patent family members are listed in annex.

* Special categories of cited documents :

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"&" document member of the same patent family

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INTERNATIONAL SEARCH REPORT

Information on patent family members

International Application No

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Patent document cited in search report	Publication date	Patent family member(s)	Publication date
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US 5559065 A	24-09-1996	DE 4221011 A DE 59305516 D EP 0576944 A	05-01-1994 03-04-1997 05-01-1994
US 4206135 A	03-06-1980	NONE	
US 4738948 A	19-04-1988	EP 0319625 A AU 595967 B AU 8250787 A US 4822824 A	14-06-1989 12-04-1990 15-06-1989 18-04-1989
US 4088671 A	09-05-1978	NONE	

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bars, a temperature of between 150°C and 350°C, the molar ratio H_2/CO in the synthesis gas being from 1:2 to 5:1.

10. The process according to claim 9, characterized in
5 that the reaction temperature is between 170°C and 300°C, the pressure between 10 and 75 bars, the ratio H_2/CO in the synthesis gas being from 1.2:1 to 2.5:1.
11. The process according to claim 10, characterized
10 in that the temperature is from 200°C to 240°C.